AMENDMENTS TO THE SPECIFICATION

Please replace the paragraph starting on page 6, line 21, as follows.

More in detail, it has been found out, in the course of tests and studies, that in the test case where the surface of the aluminum raw material was subjected to an anodic oxidization treatment to be formed with the anodic oxidation coating[[,]] in a sulfuric acid bath with an alternating current electrolysis or a direct current electrolysis, there was obtained the anodic oxidization coating provided with an innumerable number of surface opening pores of which the diameter of the majority of the surface open pores was about 10nm[[,]]. ~~and when~~ When the aluminum material having the anodic oxidation coating was then placed in a recess made in one of metal [[mold]] molds for injection molding, and the other metal mold provided with a cavity formed in a predetermined shape was closed, and molten synthetic resin was injected into the cavity to be filled therein under pressure, and after ~~cooled~~ cooling, the closed metal molds were opened, and a resultant composite product was taken out.

When a tensile strength was applied to the synthetic resin mold of the composite, the synthetic resin molding was peeled off easily from the anodic oxidation coating of the aluminum raw material [[by]] due to a small tensile strength. When the anodic oxidation coating was observed in order to study the cause thereof, it has been found that the surface open pores thereof are [[so]] too small that the molten synthetic resin can not be invaded into the pores. Now, ~~in stead~~ instead of the sulfuric acid bath, using an oxalic acid bath, a phosphoric acid bath, a sodium hydroxide bath etc., electrolysis by a direct current was carried out and there were obtained respective aluminum materials having respective anodic oxidation coatings, and for the respective aluminum raw materials, they were subjected to the foregoing injection molding using the foregoing metal mold for injection molding, so that respective composites in which the

synthetic resin moldings were coupled with the respective anodic oxidation coatings of the respective aluminum raw materials. When the tensile strength was applied to teach of the respective synthetic resin moldings thereof, the composite using the oxalic acid bath was peeled easily. However, the composite using the phosphoric acid bath and the sodium hydroxide bath, the synthetic resin molds of the respective composites were not peeled off even by a very large tensile strength. Now, after the synthetic resin molds of the respective composites were cut off from the respective anodic oxidation coatings thereof, it was observed that the innumerable pores of the respective anodic oxidation coatings were filled with solidified synthetic resin. Thus, as a result of the comparative tests, it has been found out that if molten resin is invaded into at least most of the innumerable pores, not to mention all of the pores, in the time of the molding process, molten resin is invaded into these pores, and as a result of the solidification, ~~there is produced~~ a composite is produced in which the synthetic resin molding is coupled strongly with the aluminum raw material in such a condition that synthetic resin molding is intruded into the pores.

Please replace the paragraph at page 18, line 2, with the following.

~~In stead~~ Instead of the phosphoric acid bath, using a sodium hydroxide bath, an anodic oxidation treatment was applied to the same aluminum raw material plate as that used in Example 1. Namely, as an electrolytic bath, that of a 0.2 mol aqueous solution of sodium hydride and with a temperature in the range of about 18-20°C was used, and an electrolysis was carried out by a direct current process for about 20 minutes, at a voltage of 25V, at a current density of about 0.5A/dm^2, and thereby a porous anodic oxidation coating of which almost all of the open pores are about 30-50nm in diameter was formed.[[2]]